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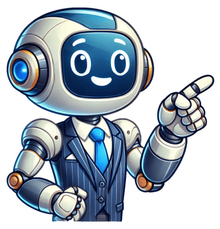

Iron(III) oxide is also known as ferric oxide. It is a chemical compound made of iron and oxygen with the formula FeO. In chemistry, we often see it as a reddish-brown powder commonly found in rust and used as a pigment. It forms when iron reacts with oxygen in the air, creating a coating that protects the iron underneath from further corrosion. Iron(III) oxide is significant in both chemistry and industry, playing a crucial role in the production of iron and steel, and its also used in magnetic materials and polishing compounds. Iron(III) oxide is also known as ferric oxide. It is a chemical compound with the formula FeO. It consists of iron and oxygen, where iron is in the +3 oxidation state. This compound commonly appears as a reddish-brown powder and is one of the main oxides of iron. Iron(III) oxide is used in various applications, including as a pigment in paints and coatings, and as a raw material in the production of iron and steel. PropertyValueFormulaFeONameIron(III) OxideAlternate NamesDiiron Trioxide, Ferric Oxide, Haematite, Hematite, Iron Sesquioxide, Oxo-(Oxoferrioxy)iron, Red Iron Oxide, Specularite The structure of Iron(III) oxide, also known as ferric oxide (FeO), consists of iron and oxygen atoms arranged in a repeating pattern. In this structure, each iron atom is surrounded by six oxygen atoms in an octahedral shape, and each oxygen atom is connected to two iron atoms. This creates a lattice structure where iron and oxygen atoms alternate, forming a stable and solid crystalline material. The arrangement allows Iron(III) oxide to have strong bonds, making it a durable compound commonly found in nature as the mineral hematite. To prepare Iron(III) oxide, you can use a simple method involving the reaction between iron and oxygen. When you heat iron in the presence of oxygen, it reacts to form Iron(III) oxide. This process is straightforward and can be performed in a lab setting. First, you take a piece of iron and heat it strongly in the presence of oxygen. The iron reacts with the oxygen to produce Iron(III) oxide, which appears as a reddish-brown powder. The chemical equation for this reaction is: 4 + 3 2FeO Alternatively, you can prepare Iron(III) oxide by decomposing iron(III) hydroxide. When you heat iron(III) hydroxide, it breaks down to form Iron(III) oxide and water. The chemical equation for this reaction is: 2() FeO + 3 Both methods result in the formation of Iron(III) oxide, commonly used in various applications, including pigments and the steel industry. PropertyDescriptionColorReddish-brownStateSolidDensity5.24 g/cmMelting Point1565CBoiling PointDecomposes before boilingSolubility in WaterInsolubleMagnetic PropertiesWeakly magnetic (paramagnetic) Iron(III) oxide reacts with acids, such as hydrochloric acid, to form iron(III) chloride and water. This reaction demonstrates its basic nature. Equation: FeO + 6 2 + 3Iron(III) oxide reacts with strong bases, like sodium hydroxide, to form sodium ferrate and water, showcasing its amphoteric nature. Equation: FeO + 2 + 3 2 + 3When heated with carbon, iron(III) oxide reduces to iron and carbon dioxide. This reaction is used in the extraction of iron from its ores. Equation: FeO + 3 2 + 3Iron(III) oxide reacts with hydrogen gas upon heating, forming iron and water. This reaction is another method for producing iron. Equation: FeO + 3 2 + 3 PropertyValueCAS Registry Number1309-37-1PubChem Compound ID14833SMILES Identifier[Fe+3].[Fe+3].[O-2].[O-2].[O-2]EU Number215-275-4Gmelin Number-81RTECS NumberNO7400000MDL NumberMFCD00011008 Iron(III) oxide serves as a pigment in paints, providing a reddish-brown color. It is widely used in art and industrial applications. Manufacturers use Iron(III) oxide as a polishing agent, especially in the final polishing of metallic jewelry and lenses, due to its abrasive properties. In steel manufacturing, Iron(III) oxide is a crucial component. It is reduced to produce iron, which is then converted into steel. Iron(III) oxide is used in magnetic storage media, such as tapes and disks, due to its magnetic properties, which help store data effectively. Industries use Iron(III) oxide as a catalyst in various chemical reactions, such as the production of ammonia in the Haber process. Iron(III) oxide is used in ceramic glazes, providing color and helping to create different finishes on pottery and tiles. Skin Irritation: Can cause irritation upon contact with skin. Eye Irritation: Dust particles may irritate the eyes. Respiratory Issues: Inhaling dust can lead to respiratory problems. Allergic Reactions: Some people may experience allergic reactions. Gastrointestinal Distress: Ingesting large amounts can cause stomach pain and nausea. Environmental Impact: Can contribute to air and water pollution if not managed properly. Iron(III) oxide has the formula FeO because it consists of iron in the +3 oxidation state combined with oxygen atoms. Iron oxide is also known as ferric oxide or hematite in its natural mineral form. Yes, rust primarily consists of iron(III) oxide (FeO), formed when iron reacts with oxygen and moisture. Iron oxide can cause skin, eye, and respiratory irritation, but it is generally considered low in toxicity with proper handling. To make iron oxide rust, expose iron to oxygen and moisture, allowing it to undergo oxidation. Iron oxide refers to compounds of iron and oxygen, while ferric oxide specifically denotes iron(III) oxide (FeO). Yes, vinegar, containing acetic acid, can dissolve iron oxide by reacting with it and breaking it down. Iron(III) oxide has ionic bonds between iron cations (Fe) and oxygen anions (O). No, Iron(III) oxide is insoluble in water due to its strong ionic bonds and crystalline structure. FeO is considered polar due to the difference in electronegativity between iron and oxygen, creating an uneven charge distribution. Add Tone Friendly Formal Casual Instructive Professional Empathetic Humorous Serious Optimistic Neutral 10 Examples of Public speaking 20 Examples of Gas lighting What is the primary chemical formula for rust, a common form of iron oxide? Choose the correct answer Which type of iron oxide is known as magnetite? Choose the correct answer What color is typically associated with iron(III) oxide? Choose the correct answer Which process is primarily responsible for the formation of iron oxide? Choose the correct answer What is the name of the iron oxide that is commonly used as a pigment in paints and dyes? Choose the correct answer Which type of iron oxide has both ferrous and ferric iron? Choose the correct answer What is the role of iron oxide in the Earth's crust? Choose the correct answer It contributes to soil fertility It forms minerals and rocks It causes water pollution How does iron oxide affect the color of soil? Choose the correct answer It makes soil red or yellow Which of the following is NOT a common use of iron oxide? Choose the correct answer As a pigment in cosmetics In the manufacture of steel In magnetic recording media What is the process of removing iron oxide from a material called? Choose the correct answer Iron oxide is essential in chemistry and helps students understand various practical and theoretical applications related to this topic. What is Iron Oxide in Chemistry?An iron oxide refers to a group of chemical compounds made up of iron and oxygen atoms. This concept appears in chapters related to corrosion, metal reactions, and environmental chemistry, making it a foundational part of your chemistry syllabus. Molecular Formula and CompositionThe molecular formula of iron oxide can be Fe2O3 (ferric oxide), Fe3O4 (magnetite), or FeO (ferrous oxide), depending on the type. It consists of iron (Fe) and oxygen (O) atoms and is categorized under inorganic oxide compounds. Type Formula Common Name Color Category Iron(II) Oxide FeO Ferrous oxide Black Simple oxide Iron(III) Oxide Fe2O3 Ferric oxide (Hematite) Reddish brown Simple oxide Iron(II,III) Oxide Fe3O4 Magnetite Black Mixed oxide Preparation and Synthesis MethodsIron oxide can be produced naturally by the reaction of iron with oxygen in the air, forming rust.In laboratories, iron(III) oxide is prepared by heating iron in oxygen or by the controlled heating of iron compounds, such as by reacting iron salts with a base.Industrially, it can be produced from iron ore in blast furnaces. Physical Properties of Iron OxideIron oxides are generally solid, with colors ranging from red-brown (hematite) to black (magnetite).Their melting point ranges from about 1,370C (FeO) to 1,565C (Fe2O3). They are odorless, insoluble in water, but can dissolve in acids.Magnetite displays magnetic properties, while hematite is weakly magnetic. Chemical Properties and ReactionsIron oxides react with acids to form iron salts and water.They participate in redox reactions, such as the thermite reaction with aluminum, which produces molten iron.Iron oxides are also involved in corrosion and rusting reactions when iron reacts with oxygen and water.These reactions are important in environmental chemistry and metallurgy. Frequent Related ErrorsConfusing Fe2O3 (ferric oxide) with Fe3O4 (magnetite).Equating rust with pure iron oxide; rust is actually a hydrated iron(III) oxide, not just Fe2O3.Mixing up iron oxidation states: +2 in FeO, +3 in Fe2O3, both in Fe3O4. Uses of Iron Oxide in Real LifeIron oxide is widely used in pigments (in paints, cosmetics, and colored concrete), magnetic materials, and sunscreens.Black iron oxide is key in anti-corrosion paints; red and yellow iron oxides are common in makeup and food coloring.Iron oxide nanoparticles find use in medicine for imaging and targeted drug delivery.The steel industry relies on it as iron ore. Relation with Other Chemistry ConceptsIron oxide is closely related to redox reactions and chemical properties of metals, helping students bridge the gap between the periodic table and real-world chemical changes like corrosion and industrial production. Step-by-Step Reaction ExampleCarbothermal reduction in steel making:Fe2O3 + 3CO 2Fe + 3CO2Explanation:Iron(III) oxide reacts with carbon monoxide at high temperature. Iron (Fe) is reduced, and carbon monoxide is oxidized to carbon dioxide. Lab or Experimental TipsRemember iron oxide colors by the simple rule: red-brown is Fe2O3 (hematite), black is Fe3O4 (magnetite). Vedantu educators recommend observing magnetic propertiesonly magnetite will stick to a magnet in the lab! Try This YourselfWrite the IUPAC name of Fe2O3.Identify which oxide is magnetic: FeO, Fe2O3, or Fe3O4.Give two real-life examples using iron oxide pigments. Final Wrap-UpWe explored iron oxideits structure, properties, reactions, and real-life importance. If you want more detailed step-by-step explanations, jump into live sessions and study notes at Vedantu.Learn about related redox reactionson Vedantus Chemistry library.Other Names for this SubstanceIron oxide (Fe2O3)Caput mortuumColloidal ferric oxideDeanoxFerrugoDeleted or Replaced CAS Registry Numbers1343-09-5, 8011-97-0, 8049-50-1, 12000-93-0, 12002-17-4, 12227-87-1, 60880-86-6, 65455-44-9, 65637-71-0, 88528-26-1, 90452-21-4, 110736-41-9, 118277-31-9, 129131-59-5, 131874-41-4, 135507-53-8, 147229-90-1, 147229-91-2, 160186-10-7, 177715-24-1, 185464-44-2, 188357-78-0, 220787-06-4, 253310-52-0, 448923-71-5, 741267-31-2, 1115688-11-3, 1146982-11-7, 1210992-56-5, 1382787-02-1, 1397708-80-3, 1430053-95-4, 2001597-72-2, 2361022-24-2, 2564724-57-6, 2808427-72-5, 2923745-42-8, 2925159-84-6, 2959468-22-3 Iron (III) Oxide consists of the elements Iron and Oxygen. Iron is a metal present in group 8 of the periodic table. Its atomic number is 26 and is represented by the symbol Fe. It is known as the most used and cheapest metal as it is the chief constituent of the earth's core. Oxygen is highly reactive non=metal. It is present in the chalcogen group in the periodic table. Its atomic number is 8 and is represented by the symbol O. Oxygen is an oxidizing agent and forms oxides with many elements. What is Iron (III) Oxide?Iron (III) Oxide is an inorganic compound and one of the most common oxides. Its chemical formula is Fe2O3. In Latin, Iron is called a Ferrum. The valency of the Iron is 2 and 3. When it combines with valency 2, It is called Ferrous and when it combines with valency 3, It is called Ferric. Iron (III) Oxide occurs naturally in rocks of all ages. IUPAC name of Iron (III) Oxide is Ferric Oxide. It is also called Hematite or Red Iron Oxide as it appears in Red-Brown color. It is an odorless and neutral compound as the pH value is 7. Structure of Iron (III) OxideIron (III) oxide consists of 3 oxygen atoms and two Iron atoms. Here Iron is metal and Oxygen is non-metal. So bonds formed between Iron and Oxygen are called Ionic bonds. The bond formation in Iron (III) oxide depends on the difference in electronegativity between Iron and Oxygen atoms. Iron (III) Oxide StructurePreparation of Iron (III) OxideIron (III) Oxide is obtained from the oxidation of Iron. In laboratories, it is prepared by electrolyzing the sodium bicarbonate solution along with inert electrolyte & Iron anode. 4Fe + 3O2 + 2H2O 4FeO(OH) The resultant hydrated Iron (III) Oxide FeO(OH) is dehydrated at 200C to produce Ferric Oxide. 2FeO(OH) Fe2O3 + H2O Physical Properties of Iron (III) OxideChemical Formula- Fe2O3Molecular Weight- 159.69 gram/molMelting Point- 1475C to 1565CWhen boiled Iron (III) Oxide gets decomposed.Density- 5.242 g/cm3Appears in reddish-brown color and it is in solid form at room temperature.Ferric oxide is odorless.It does not conduct electricity.Chemical Properties of Iron (III) OxideIron (III) Oxide is not soluble in water but soluble in strong acid.When it is dissolved in water it forms a weak base called Ferric Hydroxide.The oxidation state of Iron in oxide is +3.Carbothermal Reduction: Reducing the Iron (III) Oxide with a reducing agent like carbon gives the metal present in oxide i.e., Iron. It is applicable to all metal oxides.Fe2O3 + 3CO 2Fe + 3CO2 Exothermic thermite reaction: The Exothermic thermite reaction between Iron (III) Oxide and Aluminium gives Iron and Aluminium oxide.2Al + Fe2O3 2Fe + Al2O3 Uses of Iron (III) OxideFerric oxide is used as feedstock in the production of Iron.Used to apply the final polish to metal jewelry and as a pigment.Acts as an important ingredient in calamine lotion.Used as a pigment.Used in magnetic disks and tapes.It is used in cosmetics.Also used in Dental composites.Sample QuestionsQuestion 1: Is Iron (III) Oxide acidic or basic? Answer: Iron (III) Oxide is an atmospheric oxide of Iron. All the oxides of Iron, Aluminum and Tin exhibits both acidic and basic nature. So Iron (III) Oxide exhibits both qualities of acidic and basic. Question 2: Explain carbothermal reduction of Ferric oxide? Answer: Reduction of Ferric oxide with reducing agents such as carbon at temperature range of several hundred degree Celsius form Iron. The chemical reaction is given by- Fe2O3 + 3CO 2Fe + 3CO2 Question 3: Find the molecular weight of Iron (III) Oxide? Answer: Atomic weight of 1 Iron atom=55.845 Atomic weight of 1 oxygen atom=15.999 Molecular weight of Iron (III) Oxide (Fe2O3)=(255.845)+(315.999) =159.68 Molecular weight of Iron (III) Oxide is 159.68 Question 4: What is the result of an Exothermic thermite reaction with Aluminum? Answer: The result of Exothermic thermite reaction with Aluminum gives Iron and Aluminum Oxide. The chemical reaction is given as- 2Al + Fe2O3 2Fe + Al2O3 Question 5: State the general uses of Iron (III) Oxide? Answer: Iron (III) Oxide is used in polishing the glass, metals, diamonds. It is also used in paints, inks, rubber products and also in glass as a UV blocking agent. In medical field, Ferric oxide is used in dental composites. Ferric oxide, also called by its IUPAC name iron Trihydrate or iron(III) oxide, is an inorganic compound represented by the chemical formula Fe2O3 or Fe2H6O3 [1, 2]. It is one of the most important oxides of iron, the other two being ferrosoferric oxide (Fe3O4) and ferrous oxide (FeO) respectively [3, 4]. It naturally occurs as the mineral hematite [1].CAS Number1309-37-1 [2]PubChem CID56841934 [2]ChemSpider ID452497 [1]ChEBICHEBI:50819 [5]EC Number616-935-8 [2]RTECS NumberNO7400000 [6]InChI KeyYOBAEOGBNPPUQV-UHFFFAOYSA-N [2] Ferric Oxide Formula It is a reddish-brown gelatinous compound produced upon addition of alkali to solutions containing Fe(III) salts. It is also called hydrous ferric oxide and can be chemically represented either as Fe2O3.H2O or Fe(O)OH [12]. Some common forms of the hydrated ferric oxide include red lepidocrocite that occurs externally in rusticles and orange goethite that occurs on the inside of rusticles.How is Ferric Oxide preparedFerric oxide is synthesized in the laboratory through electrolysis of sodium bicarbonate solution, which acts as an inert electrolyte, and an iron anode. The electrolytic conversion of iron into hydrated iron (III) oxide is represented as: 4Fe + 3O2 + 2H2O 4 FeO(OH)The hydrated iron (III) oxide, thus produced, undergoes dehydration at around 200 C to form ferric oxide [7].2 FeO(OH) Fe2O3 + H2OFerric Iron Oxide PowderChemical Reactions with Other CompoundsOxidation of Ferrous Sulfate to Ferric SulfateIt involves oxidation of ferrous disulfide (FeS2) to ferrous sulfate (FeSO4), which is further oxidized to ferric sulfate [13]:FeS2 + 7O+ H2O FeSO4 + H2SO42FeSO4 + H2SO4 + O Fe2(SO4)3 + H2OAluminum replaces Iron from Ferric OxideAluminum and ferric oxide undergo a redox reaction by which the metal replaces iron to form aluminum oxide:Fe2O3 + 2Al Al2O3 + 2FeProperties and Characteristics of Ferric OxideMolar Mass/Molecular Weight159.68820 g/mol [1, 5]Color and AppearanceReddish-brown powder [1, 8]Melting Point1538-1566 C, 2800-2850 F [8, 9]Boiling PointN/A [8]Density5.24 g cm-3 [8]State of matter at room temperature (solid/liquid/gas)Solid [1]SolubilitySoluble in warm HCl, slightly soluble in H2SO4 [9]Solubility in WaterInsoluble [9]Magnetic Susceptibility ()3.586 X 10-3 emu/Oe.g [10]Heat Capacity103.9 J/mol.KCrystal StructureRhombohedral (form), cubic (and forms), orthorhombic (form) [11]Structure of Ferric OxideUsesIn iron industries for producing steel and alloys [9].Ferric oxide powder, also called jewelers rouge, is used for polishing lenses and metallic jewelry [14].Its granular form is used as a filtration media for pulling out phosphates in saltwater aquariums [14].As FDA-approved Pigment Brown 6 and Pigment Red 101, for use in cosmetics [14].In biomedical applications, because its nanoparticles are non-toxic and biocompatible [14].Is it DangerousProlonged exposure to its fume or dust can result in pneumoconiosis with fever, chills, aches, shortness of breath, and cough [15]. Repeated contact with eyes can discolor it and cause permanent iron staining [15]. Names Names and synonymsIron(III) oxide [Wiki] hematite[Wiki]Database IDs Please enable Javascript in order to use PubChem website. Iron oxide, often known as ferric oxide, is an inorganic substance. The Iron Oxide Formula is written as Fe2O3. It is one of the three primary iron oxides, the other two being the uncommon iron (II, and III) oxide (Fe3O4) and iron(II) oxide (FeO), which occurs naturally as the mineral magnetite. Hematite, often known as the mineral Fe2O3, is the main source of iron for the steel industry and is easily affected by acids. Iron oxide is frequently referred to as rust. Due to the same composition and shared features of rust, this term has some limited utility. Rust, which is actually hydrous ferric oxide, is classified as an ill-defined substance in Chemistry. Iron Oxide Formula Structure for Iron(III) oxideThe Iron Oxide Formula is written as Fe2O3. Iron(III) oxide, which contains two iron atoms and three oxygen atoms, is represented chemically by the compound Fe2O3. Fe2O3 has an oxidation state of 3. The difference in electronegativity between oxygen and iron affects how bonds between these two atoms form.Therefore, the Iron Oxide Formula is written as Fe2O3.Properties Of Iron (III) Oxide FormulaAn intriguing area of Chemistry is Inorganic Chemistry. The metal oxides are covered in the metallurgy chapter, where students will learn how to make and use these oxides in different industrial processes and why the Iron Oxide Formula is written as Fe2O3. Iron III oxide is one of the oxides students will learn about in this section. Given that the problems are somewhat conceptual, it is important to thoroughly research the chemical characteristics of this oxide.Oxides of Iron:One of the most prevalent oxides students have examined in prior sessions is iron III oxide. When they proceed to a senior class, students will get more detailed information on this oxide. This time, an entire chapter has been devoted to this oxide so that students may understand how it responds to various chemical and physical environments. The Iron Oxide Formula is written as Fe2O3. Inorganic Chemistry is a fascinating part of Chemistry. The metal oxides are included in the chapter related to metallurgy where students will learn how these oxides are prepared and used in various industrial processes. One of the oxides you will study in this section is Iron III oxide. The chemical properties of this oxide need to be properly studied as the questions are quite conceptual.To make it easier for students, Vedantu has developed a concept page focusing on the oxide and its physical and chemical properties. This page has been developed by the top Chemistry teachers. It addresses the basic problems a student faces while studying this particular oxide. It will also help to develop a clear concept so that students can easily clarify doubts arising in their minds on their own. You will learn what is the chemical formula of ferric oxide and its various properties to get a hold of this concept. Keep on reading and use this concept page as your revision note to ace the exams.Iron(III) Oxide: What are its Formula and Chemical Properties?Iron III oxide is one of the most common oxides you have studied in the previous classes. You will find more elaborate information related to this oxide when you advance to a higher class. This time, a section of the chapter has been dedicated to this oxide so that you can learn how it behaves in different physical and chemical conditions.The chemical formula of Iron III oxide is Fe2O3. It is obtained from a mineral ore called hematite. It is mined and then processed to obtain iron oxide for extracting the metal. Iron is also called Ferrum in Latin. The valency of this metal is 2 and 3. When it combines with valency 3, it is called Ferric. When it combines with other elements showing valency 2, it is called Ferrous. In this segment, we will study about this Iron Oxide and its chemical properties. The IUPAC name of Fe2O3 is Ferric Oxide. This oxide occurs in nature in a neutral state as it is properly balanced in terms of valence electrons.Production of Iron(III) Oxide:On oxidation of iron, Iron(III) oxide can be obtained. In laboratories, by electrolyzing sodium bicarbonate solution with an inert electrolyte and iron anode, Iron(III) oxide is produced.4Fe + 3O2 + 2H2O 4FeO(OH)The hydrated iron(III) oxide when dehydrated at 200C produces Iron(III) oxide alongside the water.2FeO(OH) Fe2O3 + H2OLet Us Check the Physical Properties of This OxideThe chemical name of Fe2O3 is Iron(III) oxide.Molecular Weight/ Molar Mass of Iron(III) oxide is 159.69 g/mol.The density of Iron(III) oxide is 5.242 g/cm.The Iron(III) oxide when boiled gets decomposed.The Melting Point of Iron(III) oxide is 1475C 1565C.In this section, you will find out how Iron III oxide is formed by chemically reacting with oxygen and forming the electrovalent bonds. The proper description of the chemical structure along with an illustrated image will help you grab hold of the concept properly. In fact, you will also find how the systematic name for Fe2O3 is given by the Chemistry experts. Understand the importance of the name and how it is assigned to this oxide for future references. By learning the method, you can also understand and discover how other oxides are named.Chemical Properties of Iron III OxideAs you all know that this is an oxide included in the Chemistry chapter related to metals, you will have to study its chemical properties elaborately. On this concept page, you will find out how the experts have perfectly segmented the properties along with good examples. This elaboration has been done using simple language so that students of all merit levels can understand it properly.Carbothermal reduction Reduction of metal oxides with a reducing agent i.e., carbon at several hundred degrees Celsius.Fe2O3 + 3CO 2Fe + 3CO2Exothermic thermite reaction with aluminum 2Al + Fe2O3 2Fe + Al2O3Let Us Take a Quick Look into the Chemical Properties of Iron III Oxide Fe2O3 .Iron III oxide is basic in nature. It forms a weak base when dissolved in water. The name of that base is Ferric Hydroxide (Fe(OH)3).Its alkaline properties can be observed when it reacts with an acid to form the respective inorganic salt and water at the end of an acid-base reaction.The oxidation state of the metal in this oxide is +3.It is not soluble in water but strongly soluble in strong acid.You will find the proper elaboration of these properties along with the iron 3 oxide balanced equation for each one of the points.Uses of Iron(III) Oxide (Fe2O3)After learning the Fe2O3 chemical name and chemical properties, you will proceed to the section where you will find its uses. It is used to manufacture die inks for stamping. It is the prime constituent used in various industries manufacturing pain, plastic, pharmaceutical products, ink and cosmetics. It is also used as a natural pigment.Iron(III) oxide is used in the production of iron as a feedstock.It can be used as a pigment.It can be used in cosmetics.It can be used in dental composites.It can be an important ingredient in calamine lotion.It can be used to apply the final polish on metallic jewelry.It can be used in magnetic disks and magnetic tapes.

What is the chemical formula of iron lll oxide. How is iron oxide formed. What is the chemical formula of iron ii oxide. What is the chemical formula of iron 2 oxide. What is the chemical formula of iron iii oxide. What is the chemical formula of iron 3 oxide. What is the chemical formula of hydrated iron oxide. What is the chemical formula of iron third oxide. What is the chemical formula of iron 111 oxide. What is the common name of iron oxide. What is the chemical formula of iron ll oxide. What's the chemical formula for iron oxide.